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NOTIFICATION

TO: ALL PRINCIPALS OF TECHNICAL SCHOOLS IN THE FET BAND AND DISTRICT HEADS OF EXAMINATIONS

**FROM: MRS P. JAPHTA
(a) CES: ASSESSMENTS INSTRUMENT DEVELOPMENT AND ITEM BANK MANAGEMENT SUBDIRECTORATE**

**SUBJECT: PHYSICAL SCIENCES P2 JUNE COMMON EXAMINATIONS
ERRATUM**

DATE: 19 JUNE 2023

The Physical Sciences P2 Grade 12 for June Examinations 2023 was written on Monday, 12 June 2023. We were made aware of errors and omissions that was discovered during the marking process.

The amendment with regards to the marking was prepared in conjunction with the examiner and the moderator of the paper. This amendment addresses the errors and omissions and also ensures that learners are not disadvantaged. The following standardised approach to marking must be adopted across the Province.

ERRATA: JUNE 2023 PHYSICAL SCIENCES P2 (MARKING GUIDELINE)

QUESTION/ VRAAG 3

3.5

Marking criteria

- E has dipole-dipole forces
- F has hydrogen bonds
- Correctly compare the strength of hydrogen bonds to dipole-dipole forces
- Relate the strength of the intermolecular forces to energy involved.

Nasienkriteria

- *E het dipool-dipoolkragte*
- *F het waterstofbindings*
- *Vergelyk die sterkte van die waterstofbindings korrek aan die dipool-dipoolkragte*
- *Verwys die sterktes van intermolekulêre kragte met energie betrokke*

- **F** has hydrogen bonds ✓ (and London forces)
- **E** has dipole-dipole forces ✓ (and London forces)
- Hydrogen bonds are stronger than dipole dipole forces ✓
- More energy is needed to overcome intermolecular forces in **F** ✓
- *F* het waterstofbindings (en Londonkragte)
- *E* het dipool-dipoolkragte (en Londonkragte)
- Waterstofbinding is sterker as dipool-dipoolkragte
- Meer energie word benodig om die intermolekulêrekrage in **F** te oorkom

OR/ OF

Marking criteria

- **E** has dipole-dipole forces
- **F** has hydrogen bonds
- Correctly compare the strength of hydrogen bonds to dipole-dipole forces
- Relate the strength of the intermolecular forces to energy involved

Nasienkriteria

- *E* het dipool-dipoolkragte
- *F* het waterstofbindings
- Vergelyk die sterkte van die waterstofbindings korrek aan die dipool-dipoolkragte
- Verwys die sterktes van intermolekulêrekrage met energie betrokke

- **F** has hydrogen bonds ✓ (and London forces)
- **E** has dipole-dipole forces ✓ (and London forces)
- Dipole-dipole forces are weaker than hydrogen bonds ✓
- Less energy is needed to overcome intermolecular forces in **E** ✓
- *F* het waterstofbindings (en Londonkragte)
- *E* het dipool-dipoolkragte (en Londonkragte)
- Dipool-dipoolkragte is swakker as waterstofbinding
- Minder energie word benodig om die intermolekulêrekrage in **E** te oorkom

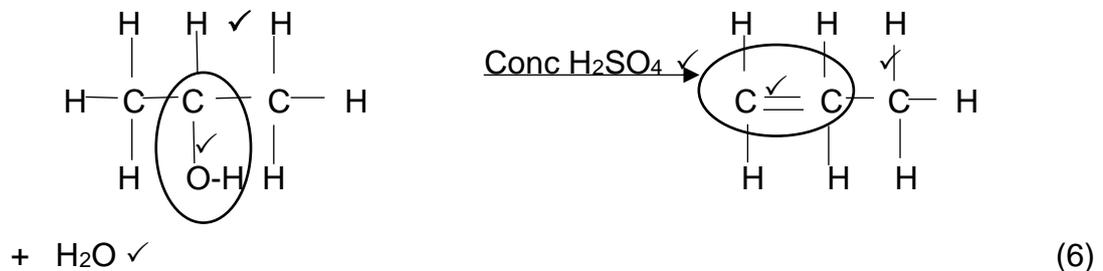
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QUESTION/ VRAAG 4

4.1.5

Marking criteria/Nasienkriteria: Organic compounds only

- Functional group/*Funksionele groep*. ✓ 1/2
- Whole structure correct/
Hele struktuur korrek ✓ 2/2



We request that this must be brought to the attention of all educators marking these papers and sincerely apologise for the inconvenience.

Yours in education.

P. Japhta

19 June 2023

Mrs P. Japhta

Date:

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